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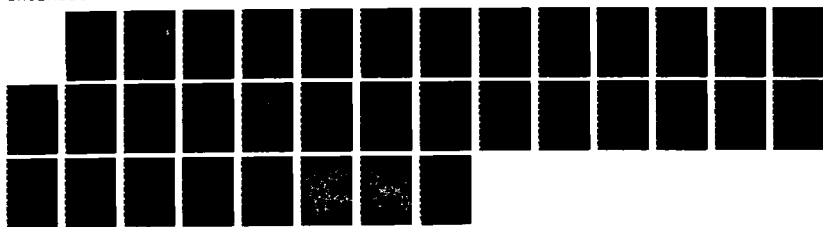
REACTIONS OF BORAZINES WITH PYRAZOLE AND SOME RELATED
STUDIES(U) KENTUCKY UNIV LEXINGTON DEPT OF CHEMISTRY
J BIELAWSKI ET AL. AUG 86 UK/DC/TR-12 N00014-84-K-0611

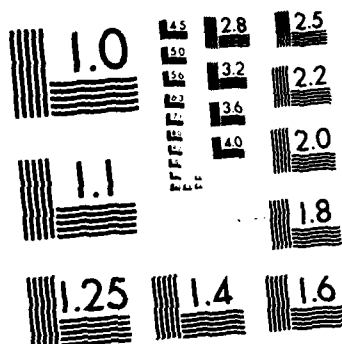
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diboron species ($R = C_2H_5$, $R' = H$) and boiling Hpz (at temperatures near 180-200 °C). The compound could be separated into conformational isomers: The lower melting one with the terminal pz groups in *cis* (equatorial) position crystallizes tetragonal in space group $I4_1/acd$; $a = b = 14.876 \text{ \AA}$, $c = 34.932 \text{ \AA}$, $Z = 16$; with the central B_2N_4 ring in boat conformation. The higher melting *trans* isomer crystallizes monoclinic in space group $P2_1/c$; $a = 8.156 \text{ \AA}$, $b = 14.720 \text{ \AA}$, $c = 7.501 \text{ \AA}$, $\beta = 95.58^\circ$, $Z = 2$, with the central B_2N_4 ring in chair conformation.

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Reactions of Borazines with Pyrazole and Some Related Studies:

μ -Amido- μ -pyrazolato Diboron Species, Novel Relatives of the Pyrazaboles

J. Bielawski⁺, M. K. Das[†], E. Hanecker[§], K. Niedenzu^{+*}, and H. Nöth[§]

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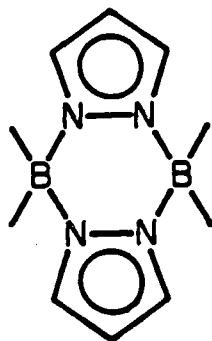
Abstract

The reaction of borazines, $(-BR-NR'-)_3$, with pyrazole, Hpz, gives μ -pyrazolato- μ -amido diboron species of the type $R(pz)B(\mu-pz)(\mu-NHR')BR(pz)$ (3) containing a central B_2N_3 ring system. Four compounds of type 3 with $R = C_2H_5$ or C_6H_5 and $R' = H$ or CH_3 , respectively, were prepared. Identical species can be obtained on interaction of bis(dimethylaminoboryl)amines of the type $R'N[BR-N(CH_3)_2]_2$ with Hpz; they are closely related to the pyrazaboles, $R_2B(\mu-pz)_2BR_2$, which contain a central B_2N_4 ring. A pathway for the formation of 3 is postulated and supported by NMR data. Conformational isomers of 3 could be identified but not separated. In contrast, $RB[N(CH_3)_2]_2$ ($R = C_2H_5$) reacts with Hpz at room temperature to give $(CH_3)_2HN-BR(pz)_2$, and in refluxing toluene to yield the pyrazabole $R(pz)B(\mu-pz)_2BR(pz)$. The latter compound was also obtained from $(-BR-NR'-)_3$ ($R' = H, CH_3$) or 3 ($R = C_2H_5, R' = H$) and boiling Hpz (at temperatures near 180-200 °C). The compound could be

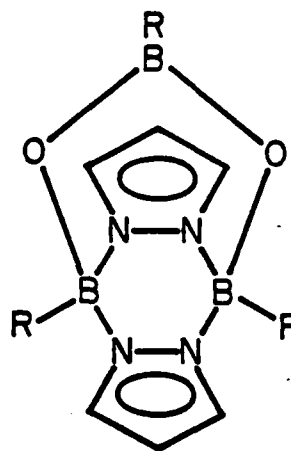
separated into conformational isomers: The lower melting one with the terminal pz groups in cis (equatorial) position crystallizes tetragonal in space group $I4_1/acd$; $a = b = 14.876 \text{ \AA}$, $c = 34.932 \text{ \AA}$, $Z = 16$; with the central B_2N_4 ring in boat conformation. The higher melting trans isomer crystallizes monoclinic in space group $P2_1/c$; $a = 8.156 \text{ \AA}$, $b = 14.720 \text{ \AA}$, $c = 7.501 \text{ \AA}$, $\beta = 95.58^\circ$, $Z = 2$; with the central B_2N_4 ring in chair conformation.

Introduction

The pyrazaboles = dimeric 1-pyrazolylboranes of the general structure 1 are heterocyclic four-coordinate boron compounds which exhibit considerable chemical and thermal stability, and a wide variety of boron- and/or carbon-substituted pyrazaboles is known.² A novel type of pyrazaboles, 2, was recently obtained by the interaction of tri-B-organylboroxins, $(-BR-O-)_3$, with pyrazole, Hpz.³



1



2

Compounds of type 2 were the first examples of neutral diboron compounds in which the two boron atoms are not only connected by two pz groups but also by a third linkage, in this case a O-BR-O unit. Thus, these compounds contain both three- and four-coordinate boron. Their seemingly unusual formation has prompted an investigation of the interaction of Hpz with other boron heterocycles; the present report describes the reaction of borazines, $(-BR-NR'-)_3$, with Hpz and some related studies.

Experimental Section

Elemental analyses were performed by the Schwarzkopf Microanalytical Laboratory, Woodside, NY. Melting points (uncorrected) were determined on a Mel-Temp block. All experiments were performed with strict exclusion of moisture.

NMR spectra were recorded on a Varian XL-200 instrument. Chemical shift data are given in ppm with positive values indicating downfield from the reference (internal Me_4Si for 1H and ^{13}C , external Et_2OBF_3 for ^{11}B); s = singlet, d = doublet, t = triplet, q = quartet, p = quintuplet, m = unresolved multiplet, and an asterisk denotes a broad signal. Coupling constants J are given in Hz. Details for HOMCOR-2D NMR experiments have been given elsewhere.⁴ Infrared spectra were recorded on a PE Model 621 or a BOMEM Model DA3 spectrometer under standard operating conditions; mass spectral data were obtained on a VG ZAB-2F spectrometer.

All solvents were dried by standard procedures. Pyrazole (= Hpz) (Aldrich Chemical Co., Milwaukee, WI) was distilled over sodium. Silazanes were obtained from Petrarch Systems, Inc., Bristol, PA; they were freshly distilled before use. Borazines, $(-BR-NR'-)_3$, were prepared by a general

method via reaction of a dihaloborane with a disilazane.⁵ The procedure is detailed below for the preparation of $(-\text{BC}_2\text{H}_5-\text{NH}-)_3$.

$(-\text{BC}_2\text{H}_5-\text{NH}-)_3$. To 223 g (2.01 mol) of $\text{C}_2\text{H}_5\text{BCl}_2$ ⁶ (cooled in an ice-ethanol bath) were added dropwise under argon cover and with stirring 366 g (2.2 mol) of $\text{HN}[\text{Si}(\text{CH}_3)_3]_2$ over a period of 4 h. The mixture was left standing overnight to reach room temperature. $(\text{CH}_3)_3\text{SiCl}$ was distilled off through a column under atmospheric pressure (bp 50-55 °C) and the residue was vacuum-distilled twice over a 25-cm silver-mantle column to yield 67.8 g (61.5%) of a main fraction of $(-\text{BC}_2\text{H}_5-\text{NH}-)_3$, bp 92-95 °C/18 torr, 56-57 °C/2 torr (lit.:⁷ bp 192-193 °C/760 torr).

NMR data (solution in CDCl_3): $\delta(^1\text{H})$ 4.68* (1 H, s), 0.95 (1 H, q, J = 6 Hz), 0.84 (3 H, t, J = 6 Hz); $\delta(^{11}\text{B})$ 36.2 (s, $h_{1/2}$ = 200 Hz).

$(-\text{BC}_2\text{H}_5-\text{NCH}_3)_3$. The compound was prepared from $\text{CH}_3\text{N}[\text{Si}(\text{CH}_3)_3]_2$ and $\text{C}_2\text{H}_5\text{BCl}_2$ in a procedure analogous to that described above for $(-\text{BC}_2\text{H}_5-\text{NH}-)_3$. Yield: 90%, bp 83-85 °C/1 torr (lit.:⁸ bp 98 °C/1.8 torr).

NMR data (solution in CDCl_3): $\delta(^1\text{H})$ 2.95 (3 H, s), 1.02 (5 H, m); $\delta(^{11}\text{B})$ 36.2 (s, $h_{1/2}$ = 220 Hz).

$(-\text{BC}_6\text{H}_5-\text{NH}-)_3$. The compound was prepared from $\text{HN}[\text{Si}(\text{CH}_3)_3]_2$ and $\text{C}_6\text{H}_5\text{BCl}_2$ ⁹ in a procedure analogous to that described above for $(-\text{BC}_2\text{H}_5-\text{NH}-)_3$ using benzene as solvent. Yield: 78%, mp 176-180 °C (lit.:¹⁰ mp 175-185 °C).

NMR data (solution in CDCl_3): $\delta(^1\text{H})$ 7.78 (2 H, m), 7.46 (3 H, m), 5.90* (1 H, s); $\delta(^{11}\text{B})$ 32.8 (s, $h_{1/2}$ = 580 Hz).

$(-BC_6H_5-NCH_3-)_3$. This compound was prepared from $C_6H_5BCl_2$ and $CH_3N[Si(CH_3)_3]_2$ in similar fashion as described above using benzene as solvent. Yield: 78%, mp 260-263 °C (lit.:¹¹ mp 247-276 °C).

NMR data (solution in $CDCl_3$): $\delta(^1H)$ 7.39 (5 H, m), 2.56 (3 H, s); $\delta(^{11}B)$ 36.3 (s, $h_{1/2}$ = 500 Hz).

$HN[BC_6H_5-N(CH_3)_2]_2$.¹² NMR data (solution in $CDCl_3$): $\delta(^1H)$ 7.40 (2 H, m), 7.25 (3 H, m), 3.70* (1 H, s), 2.75 (6 H, s); $\delta(^{11}B)$ 32.8 (s, $h_{1/2}$ = 360 Hz).

$(CH_3)_2N-B(C_6H_5)Cl$.¹³ NMR data (solution in $CDCl_3$): $\delta(^1H)$ 7.49 (2 H, m), 7.30 (3 H, m), 2.98 (3 H, s), 2.82 (3 H, s); $\delta(^{11}B)$ 36.2 (s, $h_{1/2}$ = 150 Hz).

$CH_3N[BC_6H_5-N(CH_3)_2]_2$. The compound was prepared from $(CH_3)_2N-B(C_6H_5)Cl$ and $CH_3N[Si(CH_3)_3]_2$ in a fashion analogous to the preparation of $HN[BC_6H_5-N(CH_3)_2]_2$.¹³ Yield: 41%, bp 141-143 °C/1 torr. Anal. calcd. for $C_{17}H_{25}B_2N_3$ (mol wt 293.1): C, 69.67; H, 8.60; B, 7.38; N, 14.34. Found: C, 69.13; H, 8.52; B, 7.26; N, 14.37.

NMR data (solution in $CDCl_3$): $\delta(^1H)$ 7.34 (10 H, m), 2.65 (12 H, s), 2.56 (3 H, s); $\delta(^{11}B)$ 35.6 (s, $h_{1/2}$ = 320 Hz).

The mass spectrum exhibited a strong parent ion near M/z 293; the most intense fragmentation peak was observed near M/z 248.

$(C_2H_5(pz)B(\mu-pz)(\mu-NH_2)B(C_2H_5)(pz))_3$, $R = C_2H_5$, $R' = H$. A solution of 4.1 g (25 mmol) of $(-BC_2H_5-NH-)_3$ and 5.1 g (75 mmol) of Hpz in 25 mL $CHCl_3$ was stored at room temperature for 4 weeks. Solvent was then removed under reduced pressure (15 torr) and the pasty residue was washed with 25 mL of petroleum ether. The remaining crystalline material was collected, washed

with petroleum ether and dried under vacuum to give 6.9 g (93% yield calcd for Hpz) of colorless product, mp 110-111 °C. Anal. calcd for $C_{13}H_{21}B_2N_7$ (mol wt 296.98): C, 52.58; H, 7.13; B, 7.28; N, 33.01. Found: C, 52.17; H, 7.01; B, 7.17; N, 32.81.

NMR data (solution in $CDCl_3$): $\delta(^1H)$ 7.67 (d, $J = 2.4$) + 7.63 (d, $J = 1.6$) + 7.60 (d, $J = 1.6$) + 7.56 (d, $J = 2.2$) (4 H), 7.50 (d, $J = 2.3$) + 7.34 (d, $J = 2.3$) (2 H), 6.48 (t, $J = 2.3$) + 6.43 (t, $J = 2.3$) (1 H), 6.30 (2 overlapping d) + 6.29 (2 overlapping d?) (2 H), 5.15* (2 H, s), 1.0 to 0.8 (10 H, m); $\delta(^{11}B)$ 2.7 (s, $h_{1/2} = 135$ Hz). As based on the 1H HOMCOR-2D data, the signals $\delta(^1H)$ 7.60/7.56/6.48 are assigned to the terminal pz groups and the pair 7.50/6.48 to the bridging pz group of the most abundant isomer; those at $\delta(^1H)$ 7.67/7.63/6.29 and 7.34/6.43 to the less abundant species (ratio ca. 6:1).

Mass spectral data: See text. $\nu(NH) = 3210\text{ cm}^{-1}$ (with shoulders at 3145 and 3124 cm^{-1}).

$(pz)(C_2H_5)B(\mu-pz)_2B(pz)(C_2H_5)$. A solution of 2.0 g (12 mmol) of $(-BC_2H_5-NH)_3$ and 2.5 g (37 mmol) of Hpz in 20 mL mesitylene was refluxed with stirring for 8 h. A colorless precipitate, 2.2 g, was collected and an additional 0.3 g were obtained on concentration of the filtrate to provide for a 78% yield of $(pz)(C_2H_5)B(\mu-pz)_2B(pz)(C_2H_5)$ as an isomer mixture. Isomer A, mp 152-153 °C, was obtained after two recrystallizations from cyclohexane. The cyclohexane-insoluble material was recrystallized from benzene to give a pure isomer B, mp 210-212 °C (ratio of A:B = 1:1, i.e., ca. 1 g of each product). Anal. calcd for $C_{16}H_{22}B_2N_8$ (mol wt 348.1): C, 55.20; H, 6.38; B, 6.21; N, 32.20. Found: Isomer A: C, 55.07; H, 6.47; B, 6.02; N, 31.99. Isomer B: C, 55.36; H, 6.53; B, 6.07; N, 32.26.

NMR data (solution in CDCl_3): Isomer A: $\delta(^1\text{H})$ 7.64 (2 H, d, $J = 2.5$), 7.61 (1 H, d, $J = 1.3$), 7.17 (1 H, d, $J = 2.5$), 6.52 (1 H, t, $J = 2.5$), 6.14 (1 H, unresolved t = two overlapping d), 1.34 (2 H, q, $J = 8$), 0.71 (3 H, t, $J = 8$); $\delta(^{11}\text{B})$ 2.2 (s, $h_{1/2} = 190$ Hz). Isomer B: $\delta(^1\text{H})$ 7.67 (2 H, d, $J = 2.5$), 7.64 (1 H, d, $J = 1.4$), 7.27 (1 H, d, $J = 2.2$), 6.55 (1 H, t, $J = 2.5$), 6.24 (1 H, unresolved t = two overlapping d), 0.91 (2 H, q, $J = 8$), 0.67 (3 H, t, $J = 8$); $\delta(^{11}\text{B})$ 2.2 (s, $h_{1/2} = 190$ Hz).

In the 13-eV mass spectrum of either isomer, only two major ion clusters at M/z 319 (parent ion P minus 29) and 281 (P minus 67), respectively, are observed in about 10:1 relative intensity.

Alternate Procedure. A mixture of 3.0 g (18 mmol) of $(-\text{BC}_2\text{H}_5-\text{NH}-)_3$ and 7.5 g (110 mmol) of Hpz was heated to reflux with stirring for 3 h. After standing overnight, the crude solid material was recrystallized from benzene to give B (above) and a small amount of an insoluble residue (which was discarded). After evaporation of the benzene from the filtrate, the resultant residue was recrystallized from cyclohexane to give A (above). The two isomers were obtained in A:B = 2:1 ratio, i.e., ca 4 and 2 g, respectively.

The same two isomers were obtained when $(-\text{BC}_2\text{H}_5-\text{NCH}_3-)_3$ was reacted with Hpz in a manner as described above.

Alternate Procedure. A solution of 3.2 g (25 mmol) of $\text{C}_2\text{H}_5\text{B}[\text{N}(\text{CH}_3)_2]_2^{14}$ in 15 mL of toluene was heated to reflux and a solution of 5.5 g (81 mmol) of Hpz in 55 mL of toluene was added dropwise with stirring. The mixture was refluxed for 6 h. On cooling to room temperature, 1.15 g (26.5%) of a colorless precipitate of essentially pure isomer B formed and was collected, washed with petroleum ether and dried. The filtrate was evaporated and the

residue was washed with petroleum ether (to remove excess Hpa) to give a mixture of A and B which was worked up as described above. An overall yield of about 60% of pure A and B in about 2:3 molar ratio was obtained.

Alternate Procedure. A mixture of 3 ($R = C_2H_5$, $R' = H, CH_3$) was heated with excess Hpz to 180-200 °C for 4 h. After cooling to room temperature excess Hpz was removed and the residue was worked up as described above to give an essentially quantitative yield of A and B in about 2:1 molar ratio.

Crystal data on the two isomers A and B of $(C_2H_5(pz)B(\mu-pz)_2B(C_2H_5)(pz))$ were obtained on a Syntex P3 automated diffractometer using graphite-monochromated $MoK\alpha$ radiation and standard operating techniques. Unit cell parameters were obtained from the least-squares refinement of the diffracting positions of 18 and 22 high-angle reflections, respectively. Intensity data were collected in the ω -scan technique in the range $2^\circ < 2\theta < 45^\circ$. Scattering factors for neutral atoms were taken as implemented in the SHELXTL structure solution package. The structures were solved by direct methods and the E maps revealed the positions of all non-hydrogen atoms. Parameters were refined by using the blocked-cascade method. After convergence was achieved in further refinement with anisotropic temperature factors, all hydrogen positions were revealed from a difference Fourier map. Refinement of their positions was included with use of isotropic temperature factors in the final refinement. Definitions are $R = \sum ||F_o| - |F_c|| / \sum |F_o|$ and $R = \sum (w^{1/2} ||F_o| - |F_c||) / \sum (w^{1/2} |F|)$, where $w^{-1} = \sigma^2(F_o) - g(F_o)^2$. A summary of the crystallographic data and data collection procedures is given in Table I.

Table I

$(\text{CH}_3)_2\text{HN-B}(\text{C}_2\text{H}_5)(\text{pz})_2$. A mixture of 3.2 g (25 mmol) of $\text{C}_2\text{H}_5\text{B}[\text{N}(\text{CH}_3)_2]_2$,¹⁴ 3.2 g (47 mmol) of Hpz, and 30 mL of hexane was stirred at room temperature for 24 h. The precipitate was collected, washed with hexane, and dried under vacuum to give 3.8 g (70%) of colorless crystals of the desired compound, mp 86-90 °C (with decomposition). The material can be recrystallized from hexane or cyclohexane.

NMR data (solution in CDCl_3): $\delta(^1\text{H})$ 7.69 (2 H, d, $J = 1.5$), 7.53 (2 H, d, $J = 2.2$), 6.7* (1 H, s), 6.26 (2 H, unsym t = two overlapping d), 2.39* (6 H, s; at -44 °C: unsym d, J ca. 6.5), 1.07 (2 H, q, $J = 7.5$), 0.72 (3 H, t, $J = 7.5$); $\delta(^{11}\text{B})$ 2.7 (s, $h_{1/2} = 120$ Hz).

$(\text{C}_2\text{H}_5)(\text{pz})\text{B}(\mu\text{-pz})(\mu\text{-NHCH}_3)\text{B}(\text{C}_2\text{H}_5)(\text{pz})$ (3, $\text{R} = \text{C}_2\text{H}_5$, $\text{R}' = \text{CH}_3$). A solution of 4.25 g (62.5 mmol) of Hpz and 5.2 g (25 mmol) of $(-\text{BC}_2\text{H}_5\text{-NCH}_3-)_3$ in 30 mL of CHCl_3 was stored at room temperature for 4 weeks. Solvent was removed under reduced pressure and the residue was washed with petroleum ether to give 6.0 g (93% yield) of the desired product, mp 104-108 °C. Anal. calcd for $\text{C}_{14}\text{H}_{23}\text{B}_2\text{N}_7$ (mol wt 311.01): C, 54.07; H, 7.45; B, 6.95; N, 31.53. Found: C, 53.92; H, 7.38; B, 6.93; N, 31.31.

NMR data (solution in CDCl_3): $\delta(^1\text{H})$ 7.75-7.61 (m) + 7.55 (d) + 7.35 (d) + 7.15 (d) (6 H), 6.54 (t, $J = 2.4$) + 6.38 (t, $J = 2.3$) + 6.30 (m = overlapping d?) + 6.23 (m) (3 H), 2.22 (d, $J = 6.2$) + 2.02 (d, $J = 6.3$) (ratio 1:1, 3 H), 1.2-0.5 (10 H, m); $\delta(^{11}\text{B})$ 3.0 (s, $h_{1/2} = 190$ Hz).

A very weak parent ion P minus 1 was observed in the low-voltage mass spectrum of the material. In the infrared spectrum, $\nu(\text{NH}) = 3152 \text{ cm}^{-1}$ (with shoulders at 3208 and 3090 cm^{-1}).

$(C_6H_5)(pz)B(\mu-pz)(\mu-NH_2)B(C_6H_5)(pz)$ (3, R = C_6H_5 , R' = H). A solution of 3.7 g (55 mmol) of Hpz on 20 mL of diethyl ether was added with stirring to a mixture of 7.4 g (27 mmol) of $HN[BC_6H_5-N(CH_3)_2]_2$ and 20 mL of ether. A colorless precipitate formed immediately and dimethylamine was evolved. The mixture was heated with stirring to reflux for 20 h. After cooling to room temperature, the precipitate was collected, washed with ether and then petroleum ether, and was dried in vacuum to give 4.3 g of product (60% yield calcd for Hpz), mp 168-172 °C. (An additional crop of much less pure material, mp 144-158 °C, can be obtained on concentration of the filtrate.) The initial material was recrystallized (with considerable losses) from cyclohexane/benzene; no change in the mp of 168-172 °C was observed. Anal. calcd for $C_{21}H_{21}B_2N_7$ (mol wt 393.07): C, 64.17; H, 5.39; B, 5.50; N, 24.94. Found: C, 63.51; H, 5.40; B, 5.54; N, 24.46.

NMR data (solution in $CDCl_3$): $\delta(^1H)$ 7.7-7.0 (16 H, series of ill-resolved multiplets with major peaks at 7.57, 7.49, 7.37, 7.24, 7.08; and minor peaks at 7.72, 7.52, 7.20, 6.98), 6.56 (1 H, t, J = 2.3 Hz), 6.23 (2 H, unsymmetrical t = two overlapping d), 5.9* (2 H, s); $\delta(^{11}B)$ 2.1 (s, $h_{1/2}$ = 220 Hz).

The major ion observed in the low-voltage mass spectrum is at M/z 248. $\nu(NH)$ = 3294/3169/3147/3096 cm^{-1} .

An identical material was obtained from the reaction of excess of $(-BC_6H_5-NH-)_3$ with Hpz as described below for the reaction $(-BC_6H_5-NCH_3-)_3$ with Hpz.

$(C_6H_5)(pz)B(\mu-pz)(\mu-NHCH_3)B(pz)(C_6H_5)$ (3, R = C_6H_5 , R' = CH_3). A slurry of 3.5 g (10 mmol) of $(-BC_6H_5-NCH_3-)_3$ and 2.1 g (31 mmol) of Hpz in 100 mL of hexane was refluxed with stirring for 12 h. The solid material (4.6 g) was

collected and washed with ether. It was recrystallized from cyclohexane/benzene to give 2.4 g (58%) of material, mp 186-191 °C. Anal. calcd for $C_{22}H_{23}B_2N_7$ (mol wt 407.10): C, 64.91; H, 5.69; B, 5.31; N, 24.08. Found: C, 64.81; H, 5.81; B, 5.56; N, 23.95.

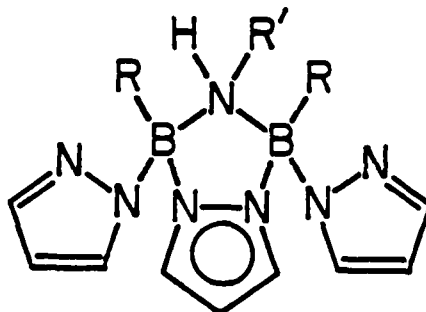
NMR data (solution in $CDCl_3$): $\delta(^1H)$ 8.42* (1 H, s), 7.8-7.0 (16 H, m with d at 7.42, 7.18, 7.08), 6.54 (1 H, 2 overlapping t of different intensity), 6.30 (2 H, m), 2.27 (d, $J = 5.6$) + 2.04 (d, $J = 6.0$) (ratio 4:1, 3 H); $\delta(^{11}B)$ 3.2 (s, $h_{1/2} = 250$ Hz). - $\nu(NH) = 3117\text{ cm}^{-1}$ (with shoulders at 3139 and 3098 cm^{-1}).

Alternate Procedure. A solution of 4.85 g (71.3 mmol) of Hpz in 20 mL of ether was added to a solution of 7.3 g (25 mmol) of $CH_3N[BC_6H_5-N(CH_3)_2]_2$ in 40 mL of ether. An immediate exothermic reaction occurred with concurrent precipitation of a colorless material. The mixture was refluxed with stirring for 25 h and the insoluble material was collected, washed with ether and then petroleum ether and was dried under vacuum to give 7.23 g (74.8% calcd for Hpz) of product, mp 190-192 °C, identical (spectroscopic data) to the material as described above.

Results and Discussion

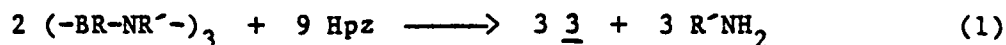
Formation and Identification of the μ -Amido- μ -pyrazolato Diboron Species
3. When $(-BC_2H_5-NH-)_3$ and Hpz (in a molar ratio not to exceed 1:4) were mixed at room temperature in $CHCl_3$ solution, a slightly exothermic reaction initiated with simultaneous generation of ammonia. The formation of a species containing four-coordinate boron was readily evidenced by the appearance of a broad ^{11}B NMR signal near 1 ppm, and signals for different pyrazolyl groups were observed in the 1H NMR spectrum (evidenced best by the pz 4-H signals).

However, even after three days at room temperature the spectra were not static and continued to change slowly. After one week minor additional changes were still observed, but there were no noteworthy additional changes after two weeks at room temperature. After four weeks the reaction had come to complete rest and no further changes in the NMR spectra were observed. At this time the solution contained (besides unreacted borazine) two isomeric species in about 6:1 molar ratio. As based on HOMCOR-2D ^1H NMR spectral data, the structure of a μ -pyrazolato- μ -amido diboron compound, $\text{R(pz)B}(\mu\text{-pz})(\mu\text{-NH}_2)\text{BR(pz)}$ (3, $\text{R} = \text{C}_2\text{H}_5$, $\text{R}' = \text{H}$), could be assigned to the isomers.



3

A mixture of the two isomers not contaminated by excess of borazine was obtained by evaporating the solvent under reduced pressure. The resultant pasty residue was then washed with petroleum ether in order to dissolve the excess of borazine and to leave the crystalline material 3 in essentially quantitative yield (calcd for Hpz). The formation of the μ -amido- μ -pyrazolato diboron species 3 can be illustrated by the following eq (1).



The more abundant isomer of 3 with $R = C_2H_5$ and $R' = H$ had the following 1H NMR data for the pz region: 7.60/7.56/6.30 ppm for the two terminal pz groups, and 7.50/6.48 ppm for the bridging pz group. The NH_2 signal for both isomers was evidenced by a broad singlet at 5.15 ppm; $\delta(^{11}B)$ for both isomers was 2.7 (s, $h_{1/2} = 135$ Hz). In conjunction with the observation of only one and reasonably sharp ^{11}B NMR signal for the material, the 1H NMR data clearly dispute the existence of a pyrazabole structure, i.e., $(H_2N)RB(\mu-pz)_2BR(pz)$. Hence, it is worth noting that compounds of this latter (isomeric) type, which are true pyrazaboles = containing a central $B(\mu-pz)_2B$ unit and a B_2N_4 ring, have previously been described as products of the interaction of Hpz with (monoboron) aminoboranes (see below).¹³

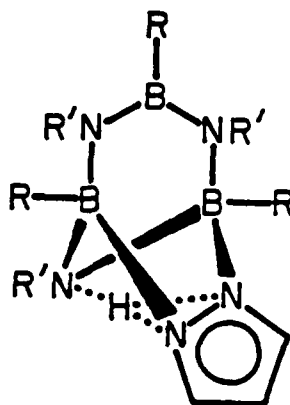
The predominant ion cluster observed in the low-voltage mass spectrum of the freshly isolated material 3 with $R = C_2H_5$ and $R' = H$ was observed in the M/z 200 region, corresponding to a composition $B_2(pz)_2(C_2H_5)(NH) = C_8H_{21}B_2N_5$, which was confirmed by the calculated and observed relative fractional abundances (f.a.):

M/z	198	199	200	201	202
f.a., calcd	0.035	0.289	0.606	0.064	0.003
f.a., obsd	0.041	0.294	0.621	0.059	0.002

A very minor peak was observed for the parent ion P minus 1 (calcd monoisotopic $P = 297$) and a fairly strong peak at M/z 68 (= Hpz). The general appearance of the mass spectral fragmentation does not follow the usual fragmentation pattern of pyrazaboles,^{15,16} thus supporting the assumed lack of a pyrazabole structure, i.e., a central $B(\mu-pz)_2B$ moiety, in the material.

The initial steps of the reaction between $(-BC_2H_5-NH-)_3$ and Hpz were best observed when the borazine was added to a large excess of Hpz. In that case, initially only a fairly broad ^{11}B NMR signal for four-coordinate boron was observed with δ 0.4 and exhibiting a shoulder at 5.5 ppm. After standing for 2 days, an additional signal near 2.5 ppm became evident. On prolonged standing, this latter shoulder emerged as the dominant and finally sole peak (δ 2.7). Similar changes were observed in the 1H NMR spectrum. For example, after mixing of the reactants, the quartet/triplet (0.95/0.84 ppm) pattern of the ethyl groups of the borazine was not observed. Rather, the mixture exhibited a broad singlet (δ 1.12) and two triplets (δ 0.80 and 0.58, respectively), the latter appearing in 2:1 ratio. On standing, the signals became more and more complex until (after about two weeks at room temperature) the multiplet of the final product emerged. Furthermore, initially the signals of unreacted Hpz at δ 7.54 (2 H, d) and 6.36 (1 H, t) as well as a signal pair δ 7.72(2 H, d)/6.30(1 H, t) were observed. Constant changes occurred until, ultimately, the signals of the final product and of Hpz remained.

It has been shown that the initial product of the interaction of equimolar amounts of a tri-B-organylboroxin, $(-BR-O-)_3$, with Hpz is an adduct containing one three-coordinate and two four-coordinate boron atoms.¹⁷ It seems reasonable to assume that a similar species, 4 ($\delta(^{11}B)$ ca. 0.4 for the four-coordinate B atoms?), is formed on interaction of $(-BR-NR'-)_3$ with Hpz. However, the third boron atom of the borazine may coordinate with the excess of Hpz to cause the observed signal $\delta(^{11}B)$ 5.5. Such an initial intermediate would account for the observation of two types of boron-bonded ethyl group signals in the 1H NMR spectrum in 2:1 ratio as described above. The singly

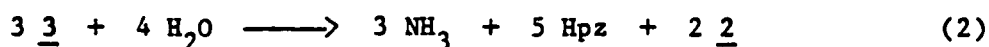


4

coordinated Hpz group may rapidly exchange with free Hpz, thus giving rise to only one common set of ^1H NMR signals for non-coordinated and mono-coordinated Hpz. The second (and much less abundant) set of pz signals as discussed above can then be assigned to the doubly coordinating Hpz group as shown in 4. Significantly, 4 contains the $\text{B}(\mu\text{-pz})(\mu\text{-NHR}')\text{B}$ skeleton of the ultimate product.

After adduct formation, the next step in the $(\text{-BR-O-})_3/\text{Hpz}$ interaction is the elimination of water and formation of two terminal B-pz bonds with retention of the bridging O-BR-O moiety to give the triply bridged pyrazabole 2.¹⁷ In the $(\text{-BR-NR'-})_3/\text{Hpz}$ case, however, the two terminal B-pz bonds are seemingly formed by attack at the B-NR'-BR-NR'-B linkages of 4 to give the observed product 3 with the elimination of amine. In this process, however, no RB(NHR')_2 is generated, since this latter should react with Hpz to give $\text{R'H}_2\text{N-BR(pz)}_2$ (see below), which was not observed.

Unfortunately, the isomers of 3 could not be separated, although it was noted that one (the less abundant) isomer seems to hydrolyze quite readily to yield the pyrazabole 2 (which also could not be separated from 3 but was clearly identified by mass spectral and NMR data). The hydrolysis 3 may proceed as is illustrated by eq (2); this would suggest that even the pyrazolyl-bridge between the two boron atoms of 3 is attacked (in order to generate a O-BR-O moiety).



Complete hydrolysis of 3 to yield 2 was achieved when the former compound was stirred with water at room temperature. The progress of the hydrolysis is readily observed by the changes in the ^1H NMR spectrum of the material. It has not been possible to separate 2 from 3; hence the preparation of 3 must be performed under strict exclusion of moisture. Also, it is difficult to completely separate an excess of Hpz from 3. Therefore it is recommended to use an excess of $(-\text{BR}-\text{NR}^--)_3$ in the preparation of 3.

The ^{11}B NMR and elemental analysis data substantiated the formation of compound 3 with $\text{R} = \text{C}_2\text{H}_5$ and $\text{R}^* = \text{CH}_3$ in the reaction of $(-\text{BC}_2\text{H}_5-\text{NCH}_3-)_3$ with Hpz. Also, the overall ^1H NMR data were in agreement with the suggested structure but, even on the basis of a ^1H HOMCOR-2D study, the spectrum could not be completely assigned. For example, despite the fact that the signal δ 6.54 is a clean 1:2:1 triplet and thus should be assigned to a bridging pz 4-H, it is coupled with two doublets (rather than the expected one) at δ 7.45 and 7.65, respectively. On the other hand, the triplet δ 6.38 is coupled only to the doublet δ 7.35. The multiplet δ 6.30 is coupled to two doublets at δ 7.15 and 7.68, respectively, but the multiplet δ 6.23 is coupled to at least

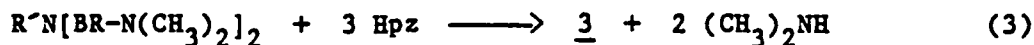
four signals in the 7.55-7.75 ppm region. Moreover, the isomer distribution (as observed by the ^1H NMR signals of the $(\text{N})\text{CH}_3$ group) differed slightly from one experiment to another although they were performed under seemingly similar conditions. Furthermore, occasionally there were indications for the formation of an additional isomer, as was shown by the appearance of another ^1H NMR signal for the methyl group. In the low-voltage mass spectrum of the material, a very weak parent ion P was observed; in the 70 eV spectrum, M/z 296 (= parent ion P minus 15) appeared instead and in addition to major fragments at M/z 282, 243, and (the base peak) 214. It must be noted, that traces of higher molecular weight species were also observed but did not recognizably affect the NMR or elemental analysis data.

The compounds were found to be thermally stable and could be sublimed under vacuum, although this procedure was not useful for purification. All attempts to promote an intermolecular condensation of 3 to yield a species of the type $\text{R}'\text{N}[\text{BR}(\mu\text{-pz})_2\text{BR}(\text{pz})]_2$ by thermal treatment failed.

The existence of species of type 3 is surprising. First of all, the reaction of tris(dimethylamino)borane with pyrazole has previously¹³ been found to give either the monoboron species $(\text{CH}_3)_2\text{HNB}(\text{pz})_3 = \text{H}[(\text{CH}_3)_2\text{NB}(\text{pz})_3]$ or true pyrazaboles with a central $\text{B}(\mu\text{-pz})_2\text{B}$ unit including the species $[(\text{CH}_3)_2\text{N}](\text{pz})\text{B}(\mu\text{-pz})_2\text{B}(\text{pz})_2$ and $[(\text{CH}_3)_2\text{N}](\text{pz})\text{B}(\mu\text{-pz})_2\text{B}[\text{N}(\text{CH}_3)_2](\text{pz})$. Similar products were obtained on reaction of $\text{C}_6\text{H}_5\text{B}[\text{N}(\text{CH}_3)_2]_2$ with Hpz . Similarly, in the present study the room temperature reaction of $\text{C}_2\text{H}_5\text{B}[\text{N}(\text{CH}_3)_2]_2$ with Hpz was found to yield $(\text{CH}_3)_2\text{HN-B}(\text{C}_2\text{H}_5)(\text{pz})_2$. In refluxing hexane, however, the reaction yielded the pyrazabole $(\text{C}_2\text{H}_5)(\text{pz})\text{B}(\mu\text{-pz})_2\text{B}(\text{C}_2\text{H}_5)[\text{N}(\text{CH}_3)_2]$ ($\delta(^{11}\text{B})$ 4.8 and 1.7), and in refluxing toluene the species $(\text{C}_2\text{H}_5)(\text{pz})\text{B}(\mu\text{-pz})_2\text{B}(\text{C}_2\text{H}_5)(\text{pz})$ was formed (see below).

Since 3 contains terminal boron-bonded pz groups that are available for coordination and, hence, could prompt the formation of a $B(\mu\text{-pz})_2B$ skeleton, there must be some inherent reason for the preferred bridging of the two boron atoms by only one pyrazolyl group and an amido group rather than two pyrazolyl groups. In this conjunction it may be noted that an analogous preferred bridging function for an $N(\text{CH}_3)_2$ group (as compared to a pz group) has recently been observed in two molybdenum poly(1-pyrazolyl)-borate complexes, where this event was thought to be due to a greater compactness of the species $B(\mu\text{-pz})(\mu\text{-NMe}_2)\text{Mo}$ as compared to $B(\mu\text{-pz})_2\text{Mo}$.¹⁸ Furthermore, the B_2N_3 ring system of 3 has also been observed in other pyrazole derivatives of boron, i.e., in the low-temperature dimerization of 1,3-dimethyl-2-(1-pyrazolyl)-diazaboracyclopentane¹⁹ and the low-temperature interaction of the latter compound with a (dimethylamino)dialkylboranes.²⁰

In the present case, the availability of a pre-existing B-N-B linkage (from the starting borazine) seems to be the prerequisite which leads to the preferential formation of 3 rather than that of a true pyrazabole. This assumption was supported by the fact that bis(dimethylaminoboryl)amines of the type $R'\text{N}[\text{BR}-\text{N}(\text{CH}_3)_2]_2$ (and containing the N-B-N-B-N arrangement) also reacted with Hpz to give the same compounds of type 3 which were obtained from the corresponding borazines. This was shown by the reaction of the two bis(dimethylaminoboryl)amines with $R = \text{C}_6\text{H}_5$ and $R' = \text{H}$ or CH_3 , respectively, as illustrated in eq (3).



Although the aromatic region from 7.7 to 7.0 ppm of the ^1H NMR spectra of the species of type 3 with $R = \text{C}_6\text{H}_5$ was quite complex, it was readily

established that two of the three pz groups of the molecules were terminal rather than bridging. For example, in the case of $R' = H$, the signal of the pz 4-H proton $\delta(^1H)$ 6.23 was not only an unsymmetrical triplet = two overlapping doublets (1 H), but irradiation of this signal led to the collapse of two doublets at 7.37 and 7.49 ppm (1 H each), respectively. On the other hand, the signal $\delta(^1H)$ 6.56 was a clear triplet (1 H), irradiation of which led to the collapse of only one doublet (2 H) at 7.57 ppm and, hence, must be assigned to a bridging pz group.

No principal differences were observed when the reaction between $(-BR-NR'-)_3$ and Hpz was carried out in the presence of a large excess of Hpz or in refluxing $CHCl_3$, benzene, or toluene rather than at room temperature. In those latter cases, the reaction was completed within several hours and there were small differences with respect to the distribution of the isomers of 3. On the other hand, minor variations with respect to the isomer distribution were also observed when the reactions were run under seemingly similar conditions at room temperature. However, in no case was the formation of a triply bridged pyrazabole of either the type $RB(\mu-pz)_2(\mu-NR'BRNR')BR$ or $RB(\mu-pz)_2(\mu-NR')BR$ observed.

Formation of Pyrazaboles. The bis(amino)borane $RB[N(CH_3)_2]_2$ ($R = C_2H_5$) reacted with excess Hpz in refluxing toluene to yield the pyrazabole $R(pz)B(\mu-pz)_2BR(pz)$. As noted above, $(-BR-NR'-)_3$ reacted with Hpz under the same conditions to give the μ -amido- μ -pyrazolato diboron compound 3. However, 3 reacted with Hpz in refluxing Hpz to form the same pyrazabole, which was also obtained directly on reaction of $(-BR-NR'-)_3$ in refluxing Hpz. Similarly, B-phenylborazines as well as compounds of type 3 with $R = C_6H_5$ were transformed to the corresponding pyrazaboles by heating in refluxing Hpz. All

of these reactions are essentially quantitative and materials of high purity were obtained quite readily.

X-ray Crystal and Molecular Structure of $R(pz)B(\mu-pz)_2BR(pz)$ ($R = C_2H_5$). The pyrazabole $(C_2H_5)(pz)B(\mu-pz)_2B(C_2H_5)(pz)$ could be separated into two isomers melting at 152-153 °C and 210-212 °C, respectively. This is the first known case that a pyrazabole of the type $RR'B(\mu-pz)_2BRR'$ could be separated into two isomers and the crystal and molecular structures of both isomers were determined by single crystal X-ray diffraction analysis.

The lower melting isomer (A) crystallizes in the tetragonal space group $I4_1/acd$ with 16 molecules in the unit cell. The central B_2N_4 ring has boat conformation and the terminal boron substituents are in cis orientation whereby the pz groups occupy the axial positions. The BN_2 plane of the boat forms an angle of 23.5° with the central N_4 plane, of which the bridging pz rings are deflected by 11.1°. The terminal pz groups are tilted by 86.3° with respect to the central N_4 plane. The structure of isomer A is illustrated in Figure 1, atomic coordinates are listed in Table II, and selected bond distances and bond angles are given in Table III.

Figure 1

Table II

Table III

The higher melting isomer (B) crystallizes in the monoclinic space group $P2_1/c$ with two molecules in the unit cell. Here, the central B_2N_4 ring has

the chair conformation (rare for a pyrazabole, see below) with the terminal boron substituents in trans orientation. The two terminal pz groups again are in the axial positions, their planes being oriented by 82.2° with respect to the N_4 plane. The N-B-N plane of the ring forms an angle of 13.6° with the N_4 plane, with which the bridging pz rings are almost coplanar. The structure of isomer B is shown in Figure 2, atomic coordinates are listed in Table IV, and selected bond distances and bond angles are given in Table III.

Figure 2

Table IV

Corresponding B-N distances of the isomers A and B are essentially identical with a surprisingly long bond to the terminal pz groups. Interestingly, the B-N bonds within the central B_2N_4 ring differ by 0.02 Å, perhaps signaling the construction of the pyrazabole skeleton by two monomeric 1-pyrazolylborane units. The B-C distance of 1.593 Å is in excellent agreement with the sum of the covalent radii.

In conjunction with previous X-ray data on pyrazaboles,²¹ some structural trends are now quite apparent. Firstly, terminal pz groups have a distinct preference for the axial positions. They are found to be orthogonal to the N_4 plane and the NN vector includes angles with the BB vector ranging from $100-120^\circ$. This arrangement seems to impair interaction with the H atoms of the bridging pz groups. Also, the boat conformation of the central B_2N_4 ring is quite predominant. The chair conformation is observed only if the two terminal substituents at each boron atom are different and trans to one another. In addition, the predominance of isomer A suggests that the boat conformation is energetically favored.

Conclusion. The present study illustrates that only the first steps of the interaction of pyrazole with boroxins or borazines, respectively, are comparable by forming adducts such as 4. Subsequently, different condensations occur to yield either 2 or 3, respectively, and this difference may be due to the electronic environment of the three-coordinate boron of the heterocycle as illustrated in 4. Still, the ready conversion of B-O bonds to B-N bonds in the $(-BR-O-)_3/Hz$ case is somewhat surprising, as is the lack of a $NR'-BR-NR'$ -bridged analog of 2. Also, it must be emphasized that formation of 3 requires a B-N-B bond in the starting material, which may be a borazine or a diborylamine. A substantial difference between 2 and 3 resides in the fact that 2 does not react with refluxing Hz . In contrast, under these same conditions 3 is readily converted to a pyrazabole of the type $R(pz)B(\mu-pz)_2BR(pz)$, which may exist in conformational isomers, as was documented for $R = C_2H_5$.

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Supplementary Material Available: Complete listing of bond angles and distances, anisotropic thermal parameter and structure factor tables for the two isomers of 4,8-diethyl-4,8-di(1-pyrazolyl)pyrazabole (19 pages). Ordering information is given on any current masthead page.

Footnotes and References

+ University of Kentucky

† On leave from Jadavpur University, Calcutta, India at the University of Kentucky.

§ University of Munich.

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Table I. Summary of Crystallographic Data and Data Collection Procedures^a

	(C ₂ H ₅)(pz)B(μ-pz) ₂ B(C ₂ H ₅)(pz)	
	isomer A	isomer B
formula	C ₁₆ H ₂₂ B ₂ N ₈	C ₁₆ H ₂₂ B ₂ N ₈
fw	335.9	335.9
cryst size, mm	0.27 x 0.30 x 0.31	0.18 x 0.34 x 0.57
cryst system	tetragonal	monoclinic
space group	I4 ₁ /acd ^b	P2 ₁ /c ^c
a, Å	14.876(6)	8.156(3)
b, Å	14.876(7)	14.720(5)
c, Å	34.932(22)	7.501(2)
α, deg	90	90
β, deg	90	95.58(3)
γ, deg	90	90
Z	16	2
V, Å ³	7730.3	896.3
d _{calcd} , g/cm ³	1.154	1.244
μ, cm ⁻¹	0.71	0.60
F(000)	2943	344
2θ, deg	2-45	2-45
scan speed, deg/min	4-29.3	4-29.3
scan width, deg	0.6	0.7
total no. of reflcns.	1607	2075
no. of unique reflcns.	997	1166
no. of variables used	159	151

Table I (contd.)

R	0.0744	0.047
R_w	0.0616	0.049
g	0.00030	0.00068
highest residual electron density, $e/\text{\AA}^3$	0.266	0.147

^aIn this and all subsequent tables esd's are given in parentheses.

^bOrigin choice 2 of International Tables; twofold axis at $1/2, 1/4, z$. The molecule shows crystallographically imposed C_2 symmetry.

^cThe molecule possesses a crystallographic inversion center at $010 = 000$.

Table II. Atomic Coordinates ($\times 10^4$) and Isotropic Thermal Parameters ($\text{\AA}^2 \times 10^3$) for Isomer A of $(\text{C}_2\text{H}_5)(\text{pz})\text{B}(\mu\text{-pz})_2\text{B}(\text{C}_2\text{H}_5)(\text{pz})$

Atom	x	y	z	U^a
N1	5411(2)	1648(2)	378(1)	45(1)
N2	5954(20)	2385(2)	352(1)	44(1)
N5	4404(2)	1526(2)	-184(1)	44(1)
N6	3604(2)	1508(2)	-372(1)	59(1)
B	4408(3)	1646(3)	258(1)	42(1)
C1	5937(3)	936(3)	451(1)	55(2)
C2	6817(3)	1207(3)	475(1)	61(2)
C3	6806(3)	2113(3)	411(1)	54(2)
C7	5083(3)	1411(2)	-436(1)	50(1)
C8	4737(3)	1319(2)	-793(1)	54(1)
C9	3822(3)	1386(3)	-738(1)	58(2)
C10	3825(3)	907(3)	471(1)	54(2)
C11	3814(3)	948(3)	905(1)	80(2)
H1	5652(22)	380(22)	478(9)	51(10)
H2	7308(27)	823(27)	500(11)	90(14)
H3	7296(22)	2576(22)	393(8)	56(10)
H7	5678(21)	1406(20)	-344(8)	43(10)
H8	5062(23)	1207(21)	-1039(8)	70(10)
H9	3362(24)	1366(22)	-909(11)	65(13)
H10a	4024(22)	325(23)	386(9)	59(11)
H10b	3177(28)	897(25)	373(10)	84(13)
H11a	4434(27)	867(27)	1004(11)	90
H11b	3473(23)	443(24)	1039(10)	90
H11c	3567(24)	1501(25)	1013(10)	90

^aEquivalent isotropic U for non-hydrogen atoms defined as one third of the trace of the orthogonalized U_{ij} tensor.

Table III. Selected Bond Distances (\AA) and Bond Angles (deg) for the cis (A) and trans (B) isomer of $(\text{C}_2\text{H}_5)(\text{pz})\text{B}(-\text{pz})_2\text{B}(\text{C}_2\text{H}_5)(\text{pz})$

Atoms	<u>A</u>	<u>B</u>
distances		
B1-N1	1.550(5)	1.577(3)
B2-N2	1.574(5)	1.557(3)
N1-N2	1.365(4)	1.356(2)
N1-C1	1.341(5)	1.334(3)
N2-C3	1.346(5)	1.340(3)
C1-C2	1.373(6)	1.366(3)
C2-C3	1.366(6)	1.362(3)
B1-C10	1.585(6)	1.597(3)
B1-N5	1.554(5)	1.548(3)
angles		
N1-B1-N4	105.8(3)	106.6(1)
B1-N1-N2	123.6(3)	126.0(2)
B2-N2-N1	123.2(3)	125.0(1)
N1-B1-C10	113.6(3)	110.7(2)
N4-B1-C10	110.6(3)	112.2(2)
N1-B1-N5	105.8(3)	107.6(1)
N4-B1-N5	108.1(3)	107.0(1)
C10-B1-N5	112.6(3)	112.5(1)

Table IV. Atomic Coordinates ($\times 10^4$) and Isotropic Thermal Parameters ($\text{\AA}^2 \times 10^3$) for Isomer B of $(\text{C}_2\text{H}_5)(\text{pz})\text{B}(\mu\text{-pz})_2\text{B}(\text{C}_2\text{H}_5)(\text{pz})$

Atom	x	y	z	U^a
N1	411(2)	10495(1)	1606(2)	34(1)
N2	-1171(2)	10216(1)	1247(2)	34(1)
N5	2824(2)	9422(1)	1642(2)	35(1)
N6	3930(2)	9637(1)	3058(2)	44(1)
B	1912(2)	10199(1)	564(3)	33(1)
C3	-2013(2)	10494(1)	2594(3)	42(1)
C2	-987(3)	10949(1)	3829(3)	45(1)
C1	518(3)	10931(1)	3173(3)	45(1)
C9	4436(3)	8846(2)	3762(3)	49(1)
C8	3690(3)	8120(2)	2831(3)	48(1)
C7	2671(2)	8510(1)	1505(3)	42(1)
C10	3083(2)	11046(1)	283(3)	42(1)
C11	2268(3)	11833(2)	-744(4)	64(1)
H1	1506(24)	11136(12)	3645(26)	50
H2	-1218(23)	11238(12)	4892(24)	50
H3	-3103(23)	10344(12)	2579(25)	50
H7	1925(24)	8270(14)	535(28)	50
H8	3837(23)	7467(16)	3091(25)	50
H9	5232(26)	8835(12)	4759(28)	50
H10a	3534(25)	11275(14)	1429(29)	50
H10b	4021(26)	10853(14)	-305(25)	50
H11a	1766(26)	11655(15)	-1881(29)	50
H11b	3016(24)	12298(15)	-958(27)	50
H11c	1351(27)	12077(14)	-206(26)	50

^aEquivalent isotropic U defined as one third of the trace of the orthogonalized U_{ij} tensor.

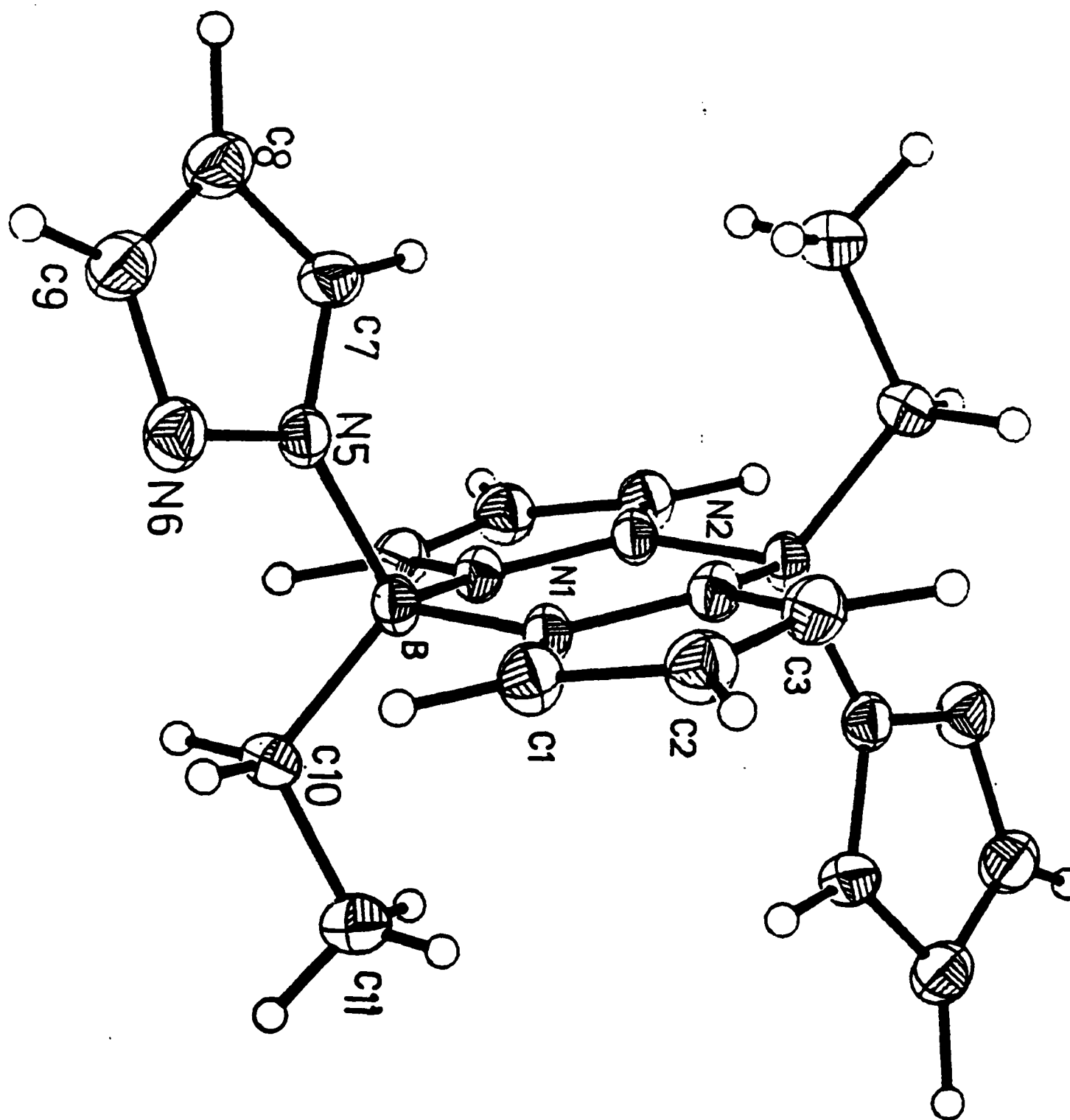


Figure 2. ORTEP plot of the isomer B of 4,8-diethyl-4,8-di(1-pyrazolyl)-pyrazabole, $(C_2H_5)(pz)B(\mu-pz)_2B(C_2H_5)(pz)$

END

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